Cambridge IGCSE

Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CANDIDATE NAME									
CENTRE NUMBER						CANDIDATE NUMBER				

CHEMISTRY **0620/33**

Paper 3 (Extended) **May/June 2015**

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.
Write in dark blue or black pen.
You may use an HB pencil for any diagrams or graphs.
Do not use staples, paper clips, glue or correction fluid.
DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.
Electronic calculators may be used.
A copy of the Periodic Table is printed on page 12.
You may lose marks if you do not show your working or if you do not use appropriate units.

At the end of the examination, fasten all your work securely together.
The number of marks is given in brackets [ ] at the end of each question or part question.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

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CAMBRIDGE International Examinations

1 Use your copy of the Periodic Table to help you answer these questions.

(a) Predict the formula of each of the following compounds.

View solution

(i) aluminium fluoride ..... [1]

(ii) arsenic oxide ..... [1]

(iii) silicon bromide ..... [1]

(b) Deduce the formula of each of the following ions.

View solution

(i) phosphide ..... [1]

(ii) barium ..... [1]

(iii) francium ..... [1]

(c) Draw a diagram showing the arrangement of the valency electrons in one molecule of the covalent compound carbon dioxide.

View solution

Use x to represent an electron from a carbon atom.
Use o to represent an electron from an oxygen atom.

[3]

[Total: 9]

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2 This question is concerned with the following oxides.

aluminium oxide
carbon monoxide
copper(II) oxide
silicon(IV) oxide
sodium oxide
sulfur dioxide
zinc oxide

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Choose **one** oxide from the above list to match each of the following descriptions. An oxide may be used once, more than once or not at all.

(a) This oxide does not react with acid or alkali. ..... [1]

(b) This oxide reacts with water to give a strong alkali solution. ..... [1]

(c) This oxide is used as a bleach. ..... [1]

(d) This oxide is amphoteric. ..... [1]

(e) This oxide has a giant covalent structure. ..... [1]

(f) This oxide is soluble in water and it is acidic. ..... [1]

[Total: 6]

View solution

3 Quicklime, which is calcium oxide, is made by heating limestone in a furnace.

$$CaCO_3(s) \rightleftharpoons CaO(s) + CO_2(g)$$

The reaction does not come to equilibrium.

(a) Suggest why the conversion to calcium oxide is complete.

View solution

..... [1]

(b) Calcium hydroxide, slaked lime, is made from calcium oxide.

View solution

Write an equation for this reaction.

..... [2]

(c) Calculate the maximum mass of calcium oxide which could be made from 12.5 tonnes of calcium carbonate. 1 tonne = 1×10^6 g.

.....

.....

..... [2]

(d) Limestone is used in agriculture to reduce the acidity of soil and for the desulfurisation of flue gases in power stations.

(i) Most crops thrive in soils whose pH is close to 7. Calcium carbonate, which is insoluble in water, and calcium oxide, which is slightly soluble in water, are both used to reduce the acidity of soils.

Suggest **two** advantages of using calcium carbonate for this purpose.

View solution

1. .....

2. ..... [2]

(ii) Explain the chemistry of desulfurisation of flue gases.

.....

.....

.....

..... [3]

(iii) Give **one** other use of calcium carbonate.

..... [1]

[Total: 11]

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4 **(a)** **(i)** Coal is a solid fossil fuel.

View solution

Name another fossil fuel.

..... [1]

(ii) Explain what is meant by the term *fossil fuel*.

View solution

.....

..... [2]

(b) The burning of fossil fuels is largely responsible for the formation of acid rain. Two of the acids in acid rain are sulfuric acid and nitric acid.

(i) Explain how the combustion of coal can form sulfuric acid.

View solution

.....

.....

..... [3]

(ii) High temperatures generated by the combustion of fossil fuels can lead to the formation of nitric acid. Explain.

.....

.....

..... [3]

(iii) Nitric acid contains nitrate ions.

Describe a test for nitrate ions.

.....

..... [2]

(iv) Explain how you could determine which one of two samples of acid rain had the higher concentration of hydrogen ions.

.....

..... [2]

[Total: 13]

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5 The law of constant composition states that all pure samples of a compound contain the same elements in the same proportion by weight.

A typical experiment to test this law is to prepare the same compound by different methods and then show that the samples have the same composition.

Methods of making copper(II) oxide include:

- heating copper carbonate,
- heating copper hydroxide,
- heating copper nitrate,
- heating copper foil in air.

(a) Complete the following equations.

View solution

(i) $CuCO_3 \rightarrow$ ..... + ... [1]

(ii) $Cu(OH)_2 \rightarrow$ ..... + ... [1]

(iii) $2Cu(NO_3)_2 \rightarrow$ ..... + $4NO_2$ + ... [2]

(b) Copper oxide can be reduced to copper by heating in hydrogen.

View solution

(i) What colour change would you observe during the reduction?

..... [1]

(ii) Explain why the copper must be allowed to cool in hydrogen before it is exposed to air.

..... [2]

(iii) Name another gas which can reduce copper(II) oxide to copper.

..... [1]

(iv) Name a solid which can reduce copper(II) oxide to copper.

..... [1]

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(c) The table below shows the results obtained by reducing the copper(II) oxide produced by different methods to copper.

View solution

(i) Complete the table.

source of copper(II) oxide	mass of copper(II) oxide / g	mass of copper / g	percentage copper / %
$CuCO_3$	2.37	1.89	79.7
$Cu(OH)_2$	2.51	1.99	
$Cu(NO_3)_2$	2.11	1.68	
Cu and O_2	2.29	1.94	

[2]

(ii) One of the samples of copper(II) oxide is impure.

Identify this sample and suggest an explanation why the percentage of copper in this sample is bigger than in the other three samples.

.....

..... [2]

[Total: 13]

View solution

6 Chemical reactions are always accompanied by an energy change.

(a) Aluminium is extracted by the electrolysis of a molten mixture which contains aluminium oxide, Al_2O_3. This decomposes to form aluminium at the negative electrode and oxygen at the positive electrode.

View solution

(i) Write an ionic equation for the reaction at the negative electrode.

..... [2]

View solution

(ii) Complete the ionic equation for the reaction at the positive electrode.

$$2O^{2-} \rightarrow \text{.....} + \text{.....}$$

[2]

View solution

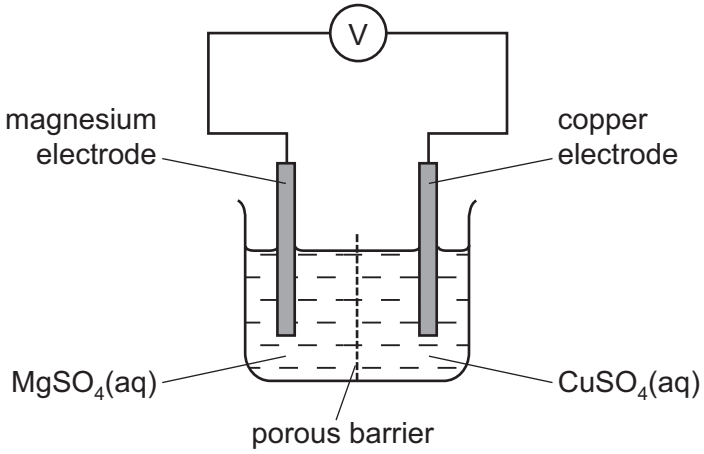
(iii) Is the reaction exothermic or endothermic? Explain your answer.

.....

..... [1]

View solution

(b) The cell shown below can be used to determine the order of reactivity of metals.



(i) Is the reaction in the cell exothermic or endothermic? Explain your answer.

.....

..... [1]

(ii) Explain why the mass of the magnesium electrode decreases and the mass of the copper electrode increases.

.....

..... [2]

(iii) How could you use this cell to determine which is the more reactive metal, magnesium or manganese?

.....

..... [2]

(c) The combustion of propane, C_3H_8, is exothermic.

View solution

Give an equation for the complete combustion of propane.

..... [2]

(d) Photosynthesis is an unusual endothermic reaction.

View solution

(i) Where does the energy for photosynthesis come from?

..... [1]

(ii) Give the word equation for photosynthesis.

..... [1]

[Total: 14]

View solution

7 **(a)** Alkanes and alkenes are both hydrocarbons.

View solution

(i) How does the structure of alkenes differ from the structure of alkanes?

..... [1]

View solution

(ii) Is the straight-chain hydrocarbon $C_{22}H_{44}$ an alkane or an alkene? Explain your choice.

.....

..... [2]

View solution

(iii) Describe how you could distinguish between pentane and pentene.

test .....

result with pentane .....

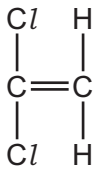
result with pentene .....

[3]

View solution

(b) Alkenes polymerise to form poly(alkenes).

(i) The alkene 1,1-dichloroethene has the structural formula given below.

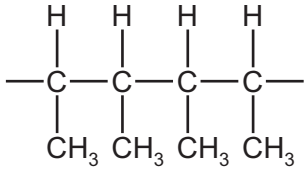


Draw the structural formula of the polymer formed by the polymerisation of 1,1-dichloroethene.

[3]

(ii) The structural formula of a different polymer is given below.

View solution



Deduce the structural formula of the monomer used to form this polymer.

[2]

(iii) There are two types of polymerisation - addition and condensation.

View solution

Explain the difference between them.

.....

.....

..... [2]

(iv) There are two types of condensation polymer.

View solution

Give the name of **one** type of condensation polymer.

.....

..... [1]

[Total: 14]

DATA SHEET
The Periodic Table of the Elements

Group																	
I	II									III	IV	V	VI	VII	0
					1 H Hydrogen 1										4 He Helium 2
7 Li Lithium 3	9 Be Beryllium 4									11 B Boron 5	12 C Carbon 6	14 N Nitrogen 7	16 O Oxygen 8	19 F Fluorine 9	20 Ne Neon 10
23 Na Sodium 11	24 Mg Magnesium 12									27 Al Aluminium 13	28 Si Silicon 14	31 P Phosphorus 15	32 S Sulfur 16	35.5 Cl Chlorine 17	40 Ar Argon 18
39 K Potassium 19	40 Ca Calcium 20	45 Sc Scandium 21	48 Ti Titanium 22	51 V Vanadium 23	52 Cr Chromium 24	55 Mn Manganese 25	56 Fe Iron 26	59 Co Cobalt 27	59 Ni Nickel 28	64 Cu Copper 29	65 Zn Zinc 30	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36
85 Rb Rubidium 37	88 Sr Strontium 38	89 Y Yttrium 39	91 Zr Zirconium 40	93 Nb Niobium 41	96 Mo Molybdenum 42	Tc Technetium 43	101 Ru Ruthenium 44	103 Rh Rhodium 45	106 Pd Palladium 46	108 Ag Silver 47	112 Cd Cadmium 48	115 In Indium 49	119 Sn Tin 50	122 Sb Antimony 51	128 Te Tellurium 52	127 I Iodine 53	131 Xe Xenon 54
133 Cs Caesium 55	137 Ba Barium 56	139 La Lanthanum 57 *	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	186 Re Rhenium 75	190 Os Osmium 76	192 Ir Iridium 77	195 Pt Platinum 78	197 Au Gold 79	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	Po Polonium 84	At Astatine 85	Rn Radon 86
Fr Francium 87	226 Ra Radium 88	227 Ac Actinium 89 †															

140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	Pm Promethium 61	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	159 Tb Terbium 65	162 Dy Dysprosium 66	165 Ho Holmium 67	167 Er Erbium 68	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71
232 Th Thorium 90	Pa Protactinium 91	238 U Uranium 92	Np Neptunium 93	Pu Plutonium 94	Am Americium 95	Cm Curium 96	Bk Berkelium 97	Cf Californium 98	Es Einsteinium 99	Fm Fermium 100	Md Mendelevium 101	No Nobelium 102	Lr Lawrencium 103

*58-71 Lanthanoid series
†90-103 Actinoid series

Key: a = relative atomic mass; X = atomic symbol; b = proton (atomic) number

The volume of one mole of any gas is 24 dm^3 at room temperature and pressure (r.t.p.).

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