Cambridge IGCSE

Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CANDIDATE NAME									
CENTRE NUMBER						CANDIDATE NUMBER			

CHEMISTRY **0620/33**

Paper 3 Theory (Core) **May/June 2018**

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.
Write in dark blue or black pen.
You may use an HB pencil for any diagrams or graphs.
Do not use staples, paper clips, glue or correction fluid.
DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.
Electronic calculators may be used.
A copy of the Periodic Table is printed on page 20.
You may lose marks if you do not show your working or if you do not use appropriate units.

At the end of the examination, fasten all your work securely together.
The number of marks is given in brackets [ ] at the end of each question or part question.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **17** printed pages and **3** blank pages.

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CAMBRIDGE International Examinations

1 The names of seven gases are given.

ammonia
ethene
helium
hydrogen
hydrogen chloride
methane
nitrogen

View solution

(a) Answer the following questions about these gases.
Each gas may be used once, more than once or not at all.
State which gas:

(i) forms an acidic solution in water

..... [1]

(ii) is monatomic View solution

..... [1]

(iii) forms 78% of clean dry air View solution

..... [1]

(iv) is the main constituent of natural gas View solution

..... [1]

(v) has molecules which each contain 14 protons. View solution

..... [1]

(b) (i) Give the names of **two** elements in the list.

View solution

..... and .....
[1]

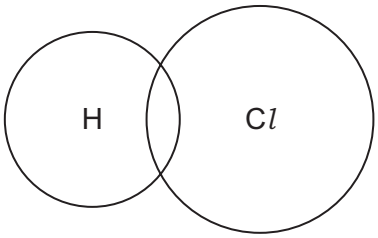
(ii) What is meant by the term *element*?

..... [1]

View solution

(c) Complete the dot-and-cross diagram to show the electron arrangement in a molecule of hydrogen chloride. Show outer electrons only.

View solution



[2]

[Total: 9]

2 **(a)** The table shows the percentage by volume of each of the gases present in the exhaust gases from a petrol engine and a diesel engine.

View solution

name	percentage by volume from a petrol engine	percentage by volume from a diesel engine
nitrogen	72.00	67.00
carbon dioxide	14.00	
water vapour	12.00	11.00
carbon monoxide	1.50	0.05
oxides of nitrogen	0.25	0.70
hydrocarbons	0.24	0.22
sulfur dioxide	0.01	0.03
oxygen	0.00	9.00
	total 100.00	total 100.00

(i) In the diesel engine, nitrogen, carbon dioxide and water vapour account for 90.00% of the exhaust gases by volume.

Calculate the percentage by volume of carbon dioxide in the exhaust gases from the diesel engine.

.....% [1]

(ii) Describe **three** ways in which the composition of the exhaust gases from the petrol engine differ from the composition of the exhaust gases from the diesel engine.

1 .....

2 .....

3 .....

[3]

(iii) Give **one** adverse effect of sulfur dioxide on health.

..... [1]

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(b) Complete the sentences about the separation of petroleum into fractions using words from the list.

View solution

boiling crystallisation density distillation kerosene poly(ethene)

Fractional ..... of petroleum produces fractions such as gasoline, diesel oil and ..... . Within each fraction, the molecules have a particular range of ..... points.

[3]

(c) Balance the chemical equation for the complete combustion of propane.

View solution

$$C_3H_8 + 5O_2 \rightarrow .....CO_2 + ...H_2O$$

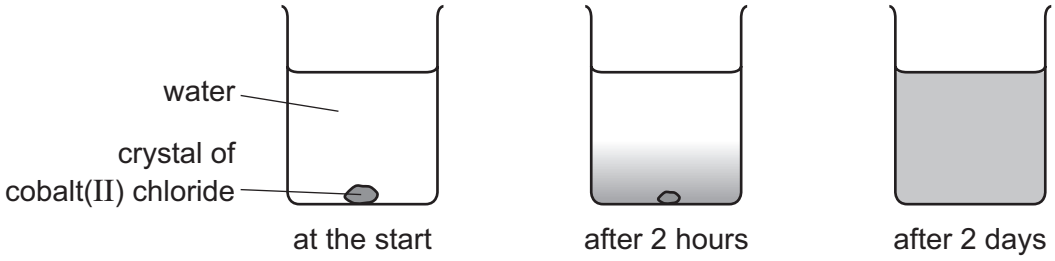
[2]

[Total: 10]

3 This question is about cobalt and its compounds.

(a) A coloured crystal of cobalt(II) chloride was placed at the bottom of a beaker containing water. After 2 days, the colour had spread throughout the water.

View solution



Explain these observations using the kinetic particle model.

.....

.....

.....

.....

..... [3]

(b) Cobalt(II) chloride can be used to test for the presence of water.

View solution

$$\underset{\substack{\text{anhydrous}\\\text{cobalt(II) chloride}}}{CoCl_2} + 6H_2O \rightleftharpoons \underset{\substack{\text{hydrated}\\\text{cobalt(II) chloride}}}{CoCl_2.6H_2O}$$

(i) What is meant by the symbol $\rightleftharpoons$?

..... [1]

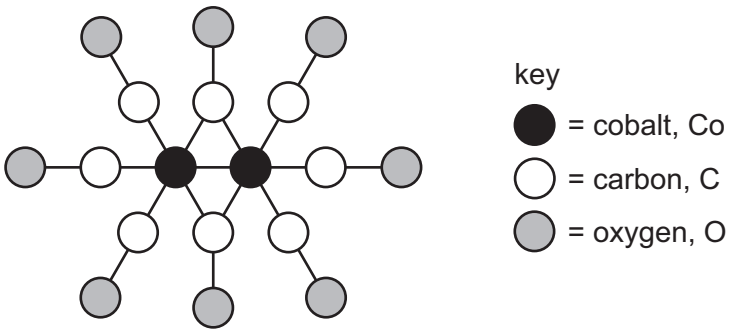
(ii) Describe how the colour of anhydrous cobalt(II) chloride changes when water is added to it.

from ..... to .....

[2]

(c) A compound of cobalt can be represented by the structure shown.

View solution



Deduce the molecular formula of this compound showing the number of cobalt, carbon and oxygen atoms.

..... [1]

(d) The table compares the reactivity of cobalt with that of three other metals.

View solution

metal	reactivity with cold water	reactivity with steam
barium	reacts rapidly	
cobalt	no reaction	reacts slowly when heated
magnesium	reacts very slowly	reacts rapidly
zinc	no reaction	reacts easily when heated

Use this information to put the metals in order of their reactivity. Put the least reactive metal first.

least reactive ⟶ most reactive

[2]

(e) Cobalt is a transition element. Lithium is an element in Group I of the Periodic Table.

View solution

Describe **three** ways in which the properties of cobalt differ from those of lithium.

1 .....

2 .....

3 .....

[2]

(f) When cobalt(II) oxide, CoO, is heated in air, an oxide with the formula Co_3O_4 is formed.

View solution

Balance the chemical equation for this reaction.

$$.....CoO + O_2 \rightarrow 2Co_3O_4$$

[1]

(g) When Co_3O_4 is heated with hydrogen, cobalt metal can be formed.

View solution

$$Co_3O_4 + 4H_2 \rightarrow 3Co + 4H_2O$$

How does this equation show that Co_3O_4 is reduced?

.....

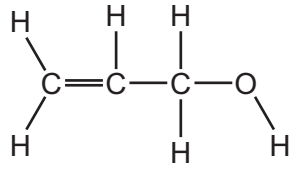
..... [1]

[Total: 13]

4 This question is about alcohols.

(a) The structure of compound **P** is shown.

View solution



(i) What feature of the structure of compound **P** shows that it is unsaturated?

..... [1]

(ii) Describe how the colour of aqueous bromine changes when an excess of compound **P** is added to it.

from ..... to .....
[2]

(iii) The melting point of compound **P** is –129 °C.
The boiling point of compound **P** is 97 °C.

What is the physical state of compound **P** at –120 °C?
Explain your answer.

.....

..... [2]

(b) Ethanol is an alcohol with two carbon atoms.

View solution

(i) Draw the structure of ethanol showing all of the atoms and all of the bonds.

[1]

(ii) Ethanol can be manufactured by fermentation.

View solution

Complete the sentences about fermentation using words from the list.

30 °C **300 °C** **distillation** **electrolysis**

elements **enzymes** **salt** **sugar**

A flask containing ..... solution and yeast is heated to ..... .

The yeast contains ..... which catalyse the reaction. The ethanol is separated by ..... . [4]

(iii) Ethanol can be used as a fuel.

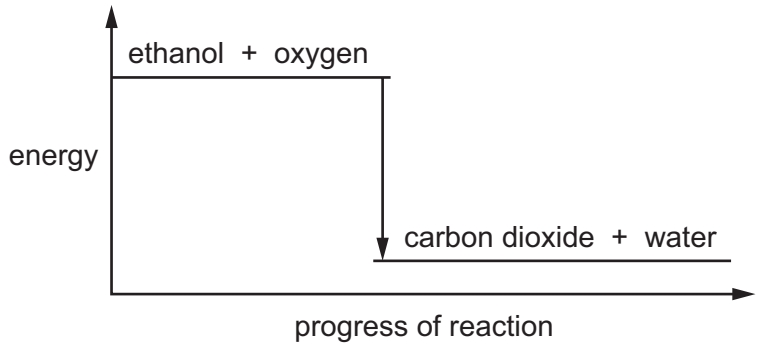
View solution

Give **one** other use of ethanol in industry.

..... [1]

(iv) The energy level diagram for the complete combustion of ethanol is shown.

View solution



Explain how this diagram shows that the reaction is exothermic.

.....

..... [1]

(v) Carbon dioxide and water can be formed when ethanol burns in a limited supply of air.

State the name of **two** other products which can be formed when ethanol burns in a limited supply of air.

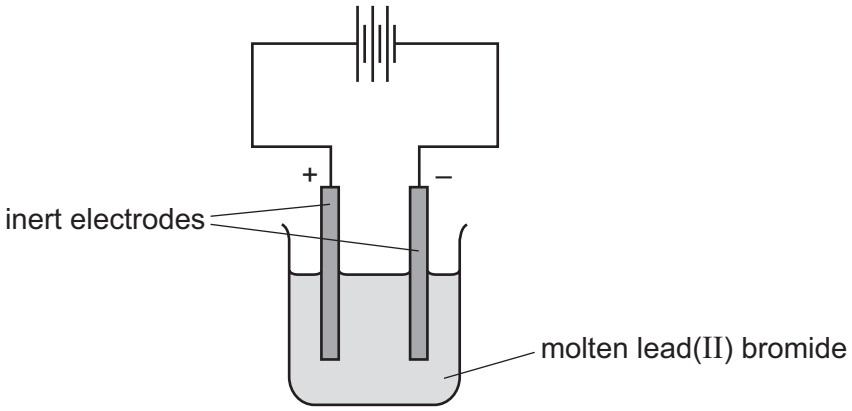
..... and .....
[2]

[Total: 14]

View solution

5 **(a)** Molten lead(II) bromide can be electrolysed using the apparatus shown.

View solution



The negative electrode is called the cathode.

State the name of the positive electrode.

..... [1]

(b) Predict the products of the electrolysis of molten lead(II) bromide at:

View solution

the negative electrode .....

the positive electrode. .....

[2]

(c) Lead(II) bromide is insoluble in cold water.
Sodium bromide is soluble in cold water.

View solution

Describe how you could obtain crystals of sodium bromide from a mixture of solid lead(II) bromide and solid sodium bromide.

.....

.....

.....

..... [3]

(d) When chlorine is bubbled through a colourless aqueous solution of sodium bromide, the solution turns orange-brown.

Name the orange-brown substance. Suggest, using ideas about reactivity of the halogens, why the reaction occurs.

orange-brown substance .....

why the reaction occurs .....

.....

[2]

(e) What is the colour of the precipitate formed when aqueous silver nitrate is added to aqueous sodium bromide?

..... [1]

(f) Complete the sentence about the formation of a chloride ion from a chlorine atom.

A chloride ion is formed when a chlorine atom gains one ..... . [1]

[Total: 10]

View solution

6 This question is about isotopes.

(a) An atom of an isotope of oxygen is represented by the symbol shown.

View solution

$^{17}_{8}O$

Describe the structure of an atom of this isotope of oxygen.
In your answer, include:

- the position of the protons, neutrons and electrons in the atom
- the number of protons, neutrons and electrons present in the atom.

.....

.....

.....

.....

.....

.....

..... [5]

(b) Which **two** statements about isotopes are correct?
Tick **two** boxes.

View solution

Isotopes of the same element have different numbers of protons.	☐
Isotopes of the same element have different numbers of neutrons.	☐
Isotopes are always radioactive.	☐
The isotope $^{235}_{92}U$ is a source of energy.	☐
$^{14}_{6}C$ and $^{14}_{7}N$ are isotopes of each other.	☐

[2]

(c) What is meant by the term *nucleon number*?

View solution

.....

..... [1]

[Total: 8]

7 This question is about Group I elements and their compounds.

(a) The properties of some Group I elements are shown in the table.

View solution

element	melting point / °C	relative hardness	observations when it reacts with cold water
lithium	181	23	
sodium	98	3	rapid bubbling but does **not** burst into flame
potassium	63		very rapid bubbling and bursts into flame
rubidium		1	bursts into flame and explodes

(i) Complete the table to estimate:
- the relative hardness of potassium
- the melting point of rubidium.

[2]

(ii) Use the information in the table to predict what you would observe when lithium reacts with cold water.

..... [1]

(b) Write the electronic structure of a sodium atom.

View solution

..... [1]

(c) When sodium reacts with water, an alkaline solution is formed.

View solution

$$2Na + 2H_2O \rightarrow 2NaOH + H_2$$

(i) Use the information in the equation to explain why the solution formed is alkaline.

..... [1]

(ii) Describe how you could use a named indicator solution to show that the solution is alkaline.

.....

..... [2]

View solution

(d) Sodium tetrahydridoborate, $NaBH_4$, is an ionic compound.

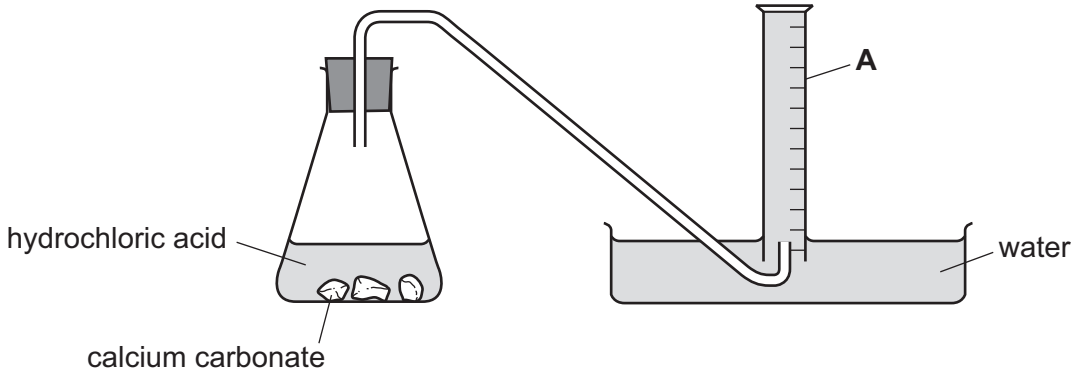
View solution

Calculate the relative formula mass of $NaBH_4$.
Show all your working.
Use your Periodic Table to help you.

relative formula mass = ... [2]

[Total: 9]

8 The apparatus shown is used to investigate the rate of reaction between calcium carbonate and hydrochloric acid at 30 °C.



(a) Name the piece of apparatus labelled **A** in the diagram.

View solution

..... [1]

(b) Describe how this apparatus can be used to find the rate of reaction.

View solution

.....

.....

..... [2]

(c) The experiment is repeated at 20 °C. All other conditions are kept the same.

View solution

How does the decrease in temperature affect the rate of reaction?

..... [1]

(d) The experiment is repeated using the same mass of different size pieces of calcium carbonate. All other conditions are kept the same.

View solution

The sizes of the pieces of calcium carbonate are:

- large pieces
- small pieces
- powder

Complete the table by writing the sizes of the pieces in the first column.

size of the pieces of calcium carbonate	initial rate of reaction in cm^3 gas / s
	5
	2
	12

[1]

(e) Give the names of the **three** products formed when calcium carbonate reacts with hydrochloric acid.

1 .....

2 .....

3 .....

[2]

[Total: 7]

View solution

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The Periodic Table of Elements

Group																	
I	II											III	IV	V	VI	VII	VIII
							1 H hydrogen 1										2 He helium 4
3 Li lithium 7	4 Be beryllium 9											5 B boron 11	6 C carbon 12	7 N nitrogen 14	8 O oxygen 16	9 F fluorine 19	10 Ne neon 20
11 Na sodium 23	12 Mg magnesium 24											13 Al aluminium 27	14 Si silicon 28	15 P phosphorus 31	16 S sulfur 32	17 Cl chlorine 35.5	18 Ar argon 40
19 K potassium 39	20 Ca calcium 40	21 Sc scandium 45	22 Ti titanium 48	23 V vanadium 51	24 Cr chromium 52	25 Mn manganese 55	26 Fe iron 56	27 Co cobalt 59	28 Ni nickel 59	29 Cu copper 64	30 Zn zinc 65	31 Ga gallium 70	32 Ge germanium 73	33 As arsenic 75	34 Se selenium 79	35 Br bromine 80	36 Kr krypton 84
37 Rb rubidium 85	38 Sr strontium 88	39 Y yttrium 89	40 Zr zirconium 91	41 Nb niobium 93	42 Mo molybdenum 96	43 Tc technetium –	44 Ru ruthenium 101	45 Rh rhodium 103	46 Pd palladium 106	47 Ag silver 108	48 Cd cadmium 112	49 In indium 115	50 Sn tin 119	51 Sb antimony 122	52 Te tellurium 128	53 I iodine 127	54 Xe xenon 131
55 Cs caesium 133	56 Ba barium 137	57–71 lanthanoids	72 Hf hafnium 178	73 Ta tantalum 181	74 W tungsten 184	75 Re rhenium 186	76 Os osmium 190	77 Ir iridium 192	78 Pt platinum 195	79 Au gold 197	80 Hg mercury 201	81 Tl thallium 204	82 Pb lead 207	83 Bi bismuth 209	84 Po polonium –	85 At astatine –	86 Rn radon –
87 Fr francium –	88 Ra radium –	89–103 actinoids	104 Rf rutherfordium –	105 Db dubnium –	106 Sg seaborgium –	107 Bh bohrium –	108 Hs hassium –	109 Mt meitnerium –	110 Ds darmstadtium –	111 Rg roentgenium –	112 Cn copernicium –		114 Fl flerovium –		116 Lv livermorium –		

Key

atomic number
atomic symbol
name
relative atomic mass

lanthanoids	57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium –	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
actinoids	89 Ac actinium –	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium –	94 Pu plutonium –	95 Am americium –	96 Cm curium –	97 Bk berkelium –	98 Cf californium –	99 Es einsteinium –	100 Fm fermium –	101 Md mendelevium –	102 No nobelium –	103 Lr lawrencium –

The volume of one mole of any gas is 24 dm^3 at room temperature and pressure (r.t.p.).