Cambridge Assessment International Education

Cambridge IGCSE™

CANDIDATE NAME									
CENTRE NUMBER						CANDIDATE NUMBER				

CHEMISTRY **0620/52**

Paper 5 Practical Test **May/June 2025**

1 hour 15 minutes

You must answer on the question paper.

You will need: The materials and apparatus listed in the confidential instructions

INSTRUCTIONS

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

INFORMATION

- The total mark for this paper is 40.
- The number of marks for each question or part question is shown in brackets [ ].
- Notes for use in qualitative analysis are provided in the question paper.

For Examiner's Use	
1	
2	
3	
Total	

This document has **12** pages. Any blank pages are indicated.

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1 You are going to investigate the temperature change when anhydrous lithium chloride dissolves in water.

Read all of the instructions carefully before starting the experiments.

Instructions
You are going to do five experiments.

Experiment 1

- Use the 50 cm^3 measuring cylinder to pour 40 cm^3 of distilled water into the 100 cm^3 beaker.
- Use the thermometer to measure the initial temperature of the water. Record this initial temperature in Table 1.1.
- Add a 2.0 g sample of anhydrous lithium chloride to the water in the beaker.
- Continually stir the mixture in the beaker using the thermometer.
- Measure the highest temperature reached by the mixture in the beaker. Record this highest temperature of the mixture in Table 1.1.
- Empty the beaker and rinse the beaker with distilled water.

Experiment 2

- Repeat Experiment 1 using 30 cm^3 of distilled water instead of 40 cm^3.

Experiment 3

- Repeat Experiment 1 using 25 cm^3 of distilled water instead of 40 cm^3.

Experiment 4

- Repeat Experiment 1 using 20 cm^3 of distilled water instead of 40 cm^3.

Experiment 5

- Repeat Experiment 1 using 15 cm^3 of distilled water instead of 40 cm^3.

(a) Complete Table 1.1.

View solution

Table 1.1

experiment	mass of anhydrous lithium chloride / g	volume of water / cm^3	initial temperature / °C	highest temperature reached / °C	temperature change / °C
1	2.0	40			
2		30			
3		25			
4		20			
5		15			

[5]

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(b) Complete a suitable scale on the *y*-axis and plot your results from Experiments 1 to 5 on Fig. 1.1.

Draw a line of best fit.

View solution

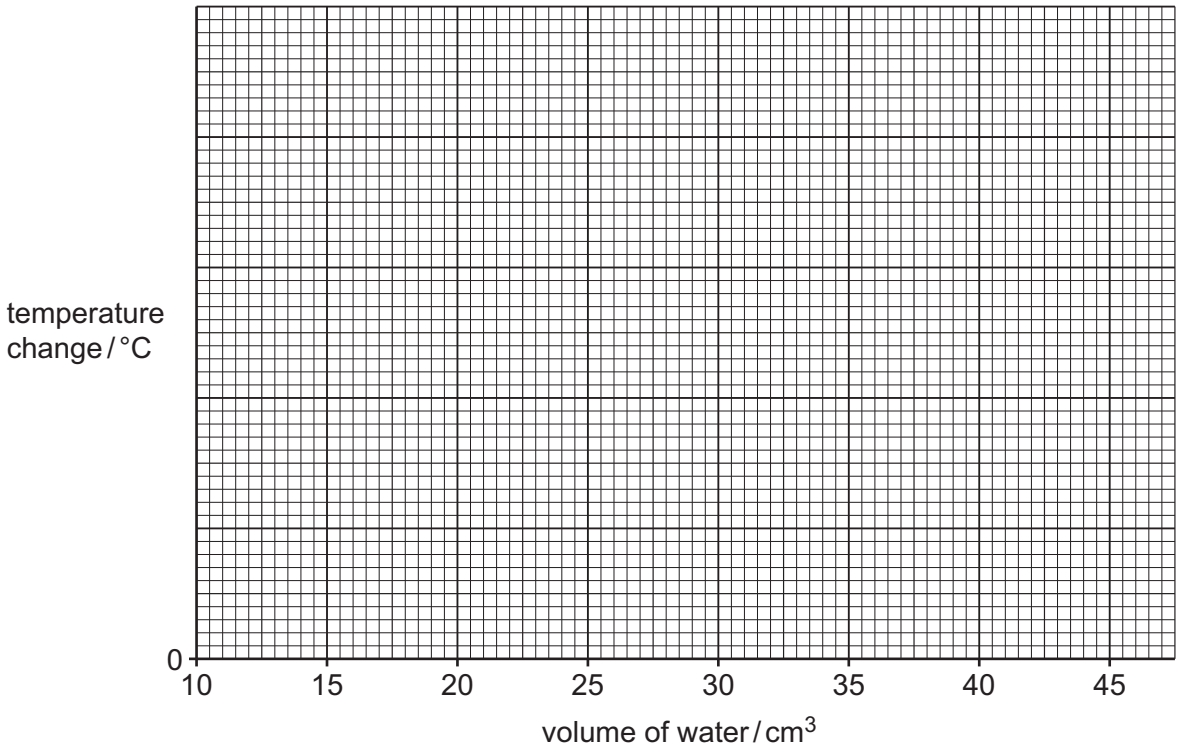

Fig. 1.1

[4]

(c) Extrapolate the line on your graph in Fig. 1.1 to deduce the temperature change when Experiment 1 is repeated with 45 cm^3 of water instead of 40 cm^3 of water.

View solution

Show clearly on Fig. 1.1 how you worked out your answer.

temperature change = ... [3]

(d) The energy, in J, given out when 2.0 g of anhydrous lithium chloride dissolves is calculated using the equation shown.

View solution

energy given out = temperature change × 4.2 × volume of water

Calculate the energy given out when 2.0 g of anhydrous lithium chloride dissolves in Experiment 5.

energy given out = ... J [1]

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(e) Estimate the temperature change when Experiment 1 is repeated using 4.0 g of anhydrous lithium chloride instead of 2.0 g.

View solution

Give a reason for your answer.

temperature change .....

reason .....

.....

[2]

(f) Explain why the results obtained would be more accurate if the beaker used in each experiment was replaced by a polystyrene cup.

.....

.....

..... [2]

(g) **(i)** Explain why using a burette instead of a measuring cylinder is an improvement.

View solution

.....

..... [1]

(ii) Explain why standing the beaker in a water-bath is **not** an improvement.

View solution

.....

..... [1]

[Total: 19]

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2 You are provided with two solids: solid **J** and solid **K**.

Do the following tests on solid **J** and solid **K**. Record all of your observations at each stage.

Tests on solid J

Add about 5 cm depth of distilled water to the boiling tube containing solid **J**. Replace the stopper in the boiling tube and shake the boiling tube to dissolve solid **J** and form solution **J**. Divide solution **J** into four approximately equal portions in four test-tubes.

(a) To the first portion of solution **J**, add aqueous sodium hydroxide dropwise and then in excess.

Record your observations.

View solution

dropwise .....

in excess .....

[2]

(b) To the second portion of solution **J**, add the sample of aqueous chlorine followed by the sample of starch solution.

Record your observations.

View solution

.....

.....

..... [2]

(c) To the third portion of solution **J**, add about 1 cm depth of dilute nitric acid followed by a few drops of aqueous silver nitrate.

Record your observations.

View solution

.....

..... [1]

(d) To the fourth portion of solution **J**, add about 1 cm depth of aqueous sodium carbonate.

Record your observations.

View solution

.....

..... [1]

(e) Identify solid **J**.

View solution

.....

..... [2]

Tests on solid K

(f) Carry out a flame test on solid **K**.

View solution

Record your observations.

..... [1]

(g) Put the remaining solid **K** into a boiling tube. Add about 3 cm depth of distilled water to the boiling tube. Put a stopper in the boiling tube and shake the boiling tube to dissolve solid **K** and form solution **K**. Divide solution **K** into two approximately equal portions in two boiling tubes.

(i) To the first portion of solution **K** in a boiling tube, add about 1 cm depth of aqueous sodium hydroxide. Warm the mixture and hold damp red litmus paper at the mouth of the boiling tube.

View solution

Record your observations.

.....

..... [1]

(ii) State the conclusion that can be made from the result of the test in **(g)(i)**.

View solution

..... [1]

(iii) To the second portion of solution **K** in a boiling tube, add about 1 cm depth of aqueous sodium hydroxide and a piece of aluminium foil. Warm the mixture and test any gas given off.

View solution

Record your observations.

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..... [2]

(h) Identify solid **K**.

View solution

.....

..... [2]

[Total: 15]

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3 Magnesium sulfate is a salt. Magnesium sulfate is soluble in water.

The solubility of a salt is the mass of the salt, in g, that dissolves in 100 cm^3 of water at a specified temperature.

Plan an investigation to determine the solubility of magnesium sulfate in water at 50 °C. Your plan must include how the solubility of magnesium sulfate, in g per 100 cm^3 of water, can be found.

You are provided with solid magnesium sulfate, distilled water and common laboratory apparatus.

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View solution

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Notes for use in qualitative analysis

Tests for anions

anion	test	test result
carbonate, CO_3^{2-}	add dilute acid, then test for carbon dioxide gas	effervescence, carbon dioxide produced
chloride, Cl^- [in solution]	acidify with dilute nitric acid, then add aqueous silver nitrate	white ppt.
bromide, Br^- [in solution]	acidify with dilute nitric acid, then add aqueous silver nitrate	cream ppt.
iodide, I^- [in solution]	acidify with dilute nitric acid, then add aqueous silver nitrate	yellow ppt.
nitrate, NO_3^- [in solution]	add aqueous sodium hydroxide, then aluminium foil; warm carefully	ammonia produced
sulfate, SO_4^{2-} [in solution]	acidify with dilute nitric acid, then add aqueous barium nitrate	white ppt.
sulfite, SO_3^{2-}	add a small volume of acidified aqueous potassium manganate(VII)	the acidified aqueous potassium manganate(VII) changes colour from purple to colourless

Tests for aqueous cations

cation	effect of aqueous sodium hydroxide	effect of aqueous ammonia
aluminium, Al^{3+}	white ppt., soluble in excess, giving a colourless solution	white ppt., insoluble in excess
ammonium, NH_4^+	ammonia produced on warming	–
calcium, Ca^{2+}	white ppt., insoluble in excess	no ppt. or very slight white ppt.
chromium(III), Cr^{3+}	green ppt., soluble in excess	green ppt., insoluble in excess
copper(II), Cu^{2+}	light blue ppt., insoluble in excess	light blue ppt., soluble in excess, giving a dark blue solution
iron(II), Fe^{2+}	green ppt., insoluble in excess, ppt. turns brown near surface on standing	green ppt., insoluble in excess, ppt. turns brown near surface on standing
iron(III), Fe^{3+}	red-brown ppt., insoluble in excess	red-brown ppt., insoluble in excess
zinc, Zn^{2+}	white ppt., soluble in excess, giving a colourless solution	white ppt., soluble in excess, giving a colourless solution

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Tests for gases

gas	test and test result
ammonia, NH_3	turns damp red litmus paper blue
carbon dioxide, CO_2	turns limewater milky
chlorine, Cl_2	bleaches damp litmus paper
hydrogen, H_2	‘pops’ with a lighted splint
oxygen, O_2	relights a glowing splint
sulfur dioxide, SO_2	turns acidified aqueous potassium manganate(VII) from purple to colourless

Flame tests for metal ions

metal ion	flame colour
lithium, Li^+	red
sodium, Na^+	yellow
potassium, K^+	lilac
calcium, Ca^{2+}	orange-red
barium, Ba^{2+}	light green
copper(II), Cu^{2+}	blue-green

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