Cambridge IGCSE

Cambridge Assessment International Education
Cambridge International General Certificate of Secondary Education

CANDIDATE NAME									
CENTRE NUMBER						CANDIDATE NUMBER			

CHEMISTRY **0620/51**

Paper 5 Practical Test **October/November 2019**

1 hour 15 minutes

Candidates answer on the Question Paper.

Additional Materials: As listed in the Confidential Instructions

READ THESE INSTRUCTIONS FIRST

Write your centre number, candidate number and name on all the work you hand in.
Write in dark blue or black pen.
You may use an HB pencil for any diagrams or graphs.
Do not use staples, paper clips, glue or correction fluid.
DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.
Electronic calculators may be used.
You may lose marks if you do not show your working or if you do not use appropriate units.
Notes for use in qualitative analysis are provided on pages 7 and 8.

At the end of the examination, fasten all your work securely together.
The number of marks is given in brackets [ ] at the end of each question or part question.

For Examiner's Use	
1	
2	
3	
Total	

This syllabus is regulated for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **8** printed pages.

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1 You are going to investigate the reaction between dilute hydrochloric acid and three different concentrations of aqueous sodium hydroxide, labelled **R**, **S** and **T**.

Read all the instructions carefully before starting the experiments.

Instructions
You are going to do three experiments.

Experiment 1

- Fill the burette up to the 0.0 cm^3 mark with dilute hydrochloric acid.
- Use the measuring cylinder to pour 20 cm^3 of solution **R** into the conical flask.
- Add six drops of methyl orange indicator to the conical flask.
- Add dilute hydrochloric acid from the burette to the conical flask, 1.0 cm^3 at a time, while swirling the conical flask, until the solution just changes colour.
- Record the burette readings in the table.
- Empty the conical flask and rinse it with distilled water.

Experiment 2

- Repeat Experiment 1 using solution **S** instead of solution **R**.
- Record the burette readings in the table.

Experiment 3

- Repeat Experiment 1 using solution **T** instead of solution **R**.
- Record the burette readings in the table.

(a) Complete the table.

View solution

burette reading / cm^3	Experiment 1 using solution **R**	Experiment 2 using solution **S**	Experiment 3 using solution **T**
final burette reading			
initial burette reading			
volume used			

[4]

(b) What colour change is observed in the conical flask at the end-point?

View solution

from ..... to ..... [2]

(c) Suggest why Universal Indicator is **not** a suitable indicator to use in these experiments.

View solution

.....

..... [1]

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(d) (i) Complete the sentences.

Experiment ..... needed the smallest volume of dilute hydrochloric acid to change the colour of the methyl orange indicator.

Experiment ..... needed the largest volume of dilute hydrochloric acid to change the colour of the methyl orange indicator.

[1]

(ii) Determine the simplest whole number ratio of volumes of dilute hydrochloric acid used in Experiments 1 and 2.

Experiment 1 ..... **:** ..... Experiment 2 [1]

(iii) Deduce the order of concentrations of the solutions of aqueous sodium hydroxide, **R**, **S** and **T**.

View solution

most concentrated .....

.....

least concentrated .....

[1]

(e) What would be the effect on the results, if any, if the solutions of aqueous sodium hydroxide were warmed before adding the dilute hydrochloric acid? Give a reason for your answer.

effect on the results .....

reason .....

View solution [2]

(f) Suggest how the reliability of the results could be checked. View solution

.....

..... [2]

(g) Suggest a different method, **not** involving an indicator, of finding the order of concentrations of the solutions of aqueous sodium hydroxide, **R**, **S** and **T**.

.....

.....

..... [3]

[Total: 17]

View solution

2 You are provided with solid **U** and liquid **V**.
Do the following tests on solid **U** and liquid **V**, recording all of your observations at each stage.

tests on solid U

(a) Do a flame test on a small sample of solid **U**.
Record your observations.

View solution

..... [1]

Add the rest of solid **U** to about 10 cm^3 of distilled water in a boiling tube. Stopper the boiling tube and shake the mixture to dissolve solid **U** and form solution **U**.

(b) Describe the colour of solution **U**.

..... [1]

Divide solution **U** into three approximately equal portions in three test-tubes.

View solution

(c) (i) Add a few drops of aqueous sodium hydroxide to the first portion of solution **U**.
Record your observations.

..... [1]

(ii) Now add an excess of aqueous sodium hydroxide to this mixture.
Record your observations.

View solution

..... [1]

(d) (i) Add a few drops of aqueous ammonia to the second portion of solution **U**.
Record your observations.

View solution

..... [2]

(ii) Now add an excess of aqueous ammonia to this mixture.
Record your observations.

View solution

..... [1]

(e) Add a few drops of dilute nitric acid and about 1 cm^3 of aqueous barium nitrate to the third portion of solution **U**. Leave to stand for 5 minutes.
Record your observations.

View solution

..... [2]

(f) Solid **U** contains three different ions.

View solution

What conclusions can you draw about the ions present in solid **U**?

.....

..... [3]

tests on liquid V

(g) Describe the appearance of liquid **V**.
Record your observations.

View solution

..... [1]

(h) Use a teat pipette to place a few drops of liquid **V** onto a watch-glass. Put the stopper back into the test-tube of liquid **V**. Use a lighted splint to touch the surface of liquid **V** carefully.
Record your observations.

View solution

..... [1]

(i) Use a spatula to transfer a small crystal of iodine carefully into the rest of liquid **V** in the stoppered test-tube. Put the stopper back into the test-tube and shake the test-tube.
Record your observations.

View solution

.....

..... [2]

(j) Draw **one** conclusion about liquid **V**.

View solution

..... [1]

[Total: 17]

3 Potassium nitrate and ammonium chloride are two salts. The energy change when they each dissolve in water is endothermic.

Plan an experiment to show which of these two salts produces the larger endothermic energy change per gram.


View solution

Your answer should include:

- any measurements you would take and record
- how the results could be used to draw a conclusion.

You are provided with potassium nitrate and ammonium chloride, distilled water and common laboratory apparatus.

.....

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.....

..... [6]

Notes for use in qualitative analysis
Tests for anions

anion	test	test result
carbonate (CO_3^{2-})	add dilute acid	effervescence, carbon dioxide produced
chloride (Cl^-) [in solution]	acidify with dilute nitric acid, then add aqueous silver nitrate	white ppt.
bromide (Br^-) [in solution]	acidify with dilute nitric acid, then add aqueous silver nitrate	cream ppt.
iodide (I^-) [in solution]	acidify with dilute nitric acid, then add aqueous silver nitrate	yellow ppt.
nitrate (NO_3^-) [in solution]	add aqueous sodium hydroxide, then aluminium foil; warm carefully	ammonia produced
sulfate (SO_4^{2-}) [in solution]	acidify, then add aqueous barium nitrate	white ppt.
sulfite (SO_3^{2-})	add dilute hydrochloric acid, warm gently and test for the presence of sulfur dioxide	sulfur dioxide produced will turn acidified aqueous potassium manganate(VII) from purple to colourless

Tests for aqueous cations

cation	effect of aqueous sodium hydroxide	effect of aqueous ammonia
aluminium (Al^{3+})	white ppt., soluble in excess, giving a colourless solution	white ppt., insoluble in excess
ammonium (NH_4^+)	ammonia produced on warming	–
calcium (Ca^{2+})	white ppt., insoluble in excess	no ppt., or very slight white ppt.
chromium(III) (Cr^{3+})	green ppt., soluble in excess	grey-green ppt., insoluble in excess
copper(II) (Cu^{2+})	light blue ppt., insoluble in excess	light blue ppt., soluble in excess, giving a dark blue solution
iron(II) (Fe^{2+})	green ppt., insoluble in excess	green ppt., insoluble in excess
iron(III) (Fe^{3+})	red-brown ppt., insoluble in excess	red-brown ppt., insoluble in excess
zinc (Zn^{2+})	white ppt., soluble in excess, giving a colourless solution	white ppt., soluble in excess, giving a colourless solution

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Tests for gases

gas	test and test result
ammonia (NH_3)	turns damp red litmus paper blue
carbon dioxide (CO_2)	turns limewater milky
chlorine (Cl_2)	bleaches damp litmus paper
hydrogen (H_2)	'pops' with a lighted splint
oxygen (O_2)	relights a glowing splint
sulfur dioxide (SO_2)	turns acidified aqueous potassium manganate(VII) from purple to colourless

Flame tests for metal ions

metal ion	flame colour
lithium (Li^+)	red
sodium (Na^+)	yellow
potassium (K^+)	lilac
copper(II) (Cu^{2+})	blue-green

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