Cambridge Assessment International Education

Cambridge IGCSE™

CANDIDATE NAME									
CENTRE NUMBER						CANDIDATE NUMBER				

CHEMISTRY **0620/52**

Paper 5 Practical Test **October/November 2020**

1 hour 15 minutes

You must answer on the question paper.

You will need: The materials and apparatus listed in the confidential instructions

INSTRUCTIONS

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

INFORMATION

- The total mark for this paper is 40.
- The number of marks for each question or part question is shown in brackets [ ].
- Notes for use in qualitative analysis are provided in the question paper.

For Examiner's Use	
1	
2	
3	
Total	

This document has **12** pages. Blank pages are indicated.

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1 You are going to investigate the rate of the reaction between sodium metabisulfite and potassium iodate.

Read all of the instructions carefully before starting the experiments.

Instructions
You are going to do five experiments. In each experiment the total volume of liquid is 45 cm^3.

Experiment 1
- Use a 10 cm^3 measuring cylinder to pour 5 cm^3 of aqueous sodium metabisulfite into the beaker.
- Use another 10 cm^3 measuring cylinder to pour 5 cm^3 of starch solution into the beaker.
- Use a 25 cm^3 measuring cylinder to pour 15 cm^3 of distilled water into the beaker.
- Use another 25 cm^3 measuring cylinder to pour 20 cm^3 of aqueous potassium iodate into the beaker. At the same time start the stop-clock.
- Stir the mixture in the beaker and continue to stir until a sudden colour change is seen.
- Stop the stop-clock and record the time in seconds to the nearest whole number in the table.
- Rinse the beaker with water.

Experiment 2
- Repeat Experiment 1 using 17 cm^3 of distilled water and 18 cm^3 of aqueous potassium iodate.

Experiment 3
- Repeat Experiment 1 using 21 cm^3 of distilled water and 14 cm^3 of aqueous potassium iodate.

Experiment 4
- Repeat Experiment 1 using 23 cm^3 of distilled water and 12 cm^3 of aqueous potassium iodate.

Experiment 5
- Repeat Experiment 1 using 25 cm^3 of distilled water and 10 cm^3 of aqueous potassium iodate.

(a) Complete the table.

View solution

experiment	1	2	3	4	5
volume of aqueous sodium metabisulfite / cm^3					
volume of distilled water / cm^3					
volume of aqueous potassium iodate / cm^3	20	18	14	12	10
time taken to change colour / s					

[4]

(b) State the sudden colour change seen.

View solution

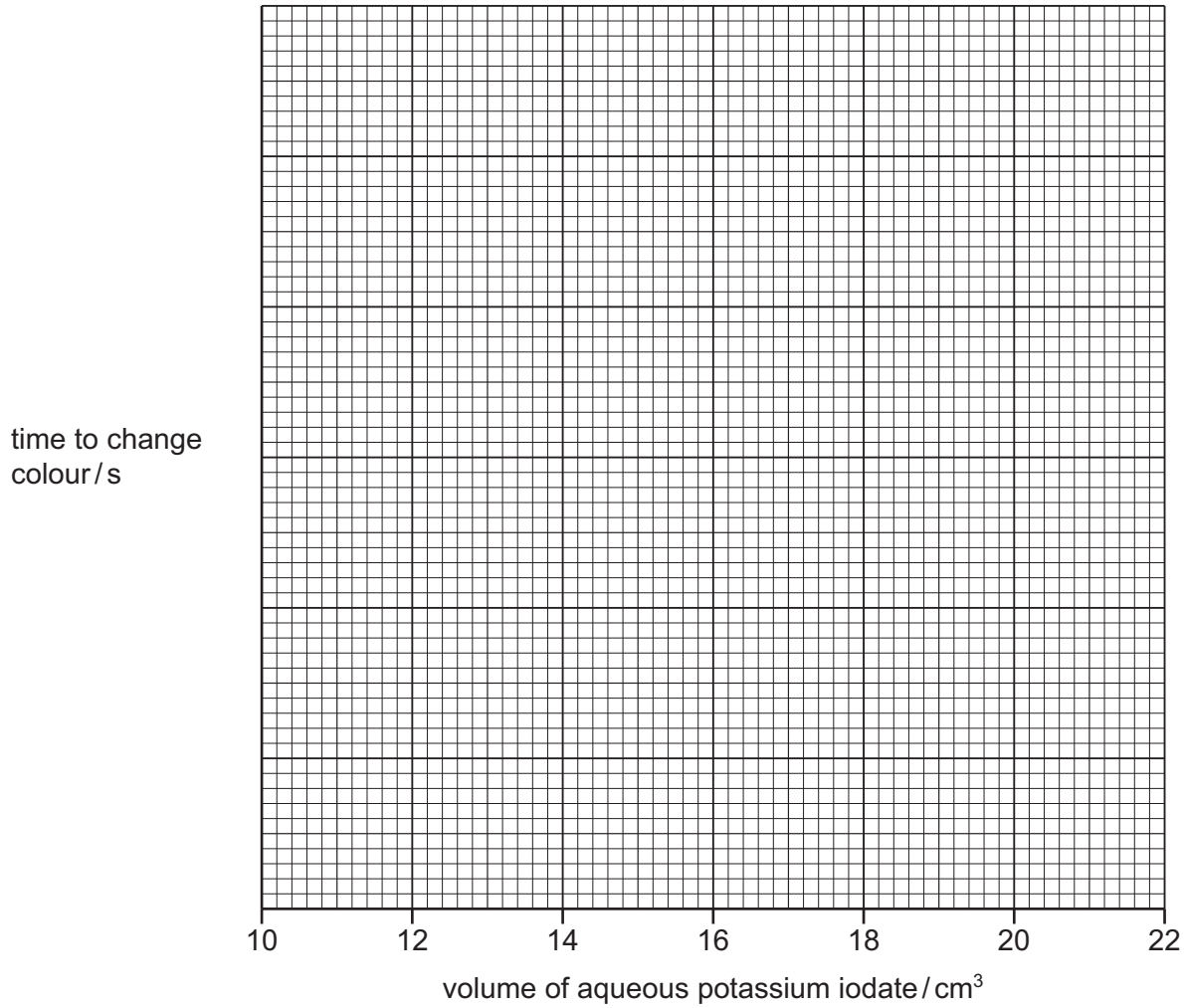
start colour .....

end colour .....

[1]

(c) Add a suitable scale for the *y*-axis and plot your results from Experiments 1 to 5 on the grid. Draw a smooth curve of best fit.

View solution



[4]

(d) (i) **From your graph**, predict the time to change colour if 16 cm^3 of aqueous potassium iodate was used.
Show clearly **on the grid** how you worked out your answer.

View solution

time to change colour = ... s [2]

(ii) Calculate the volume of distilled water required if 16 cm^3 of aqueous potassium iodate was used.

volume of distilled water = ... cm^3 [1]

(e) Sketch **on the grid** the graph you would expect if Experiments 1 to 5 were repeated at a higher temperature. [1]

View solution

(f) The concentration of potassium iodate in the reaction mixture in each experiment can be calculated using the equation shown.

View solution

$$\text{concentration} = \frac{0.05 \times \text{volume of aqueous potassium iodate}}{45}$$

(i) Calculate the concentration of potassium iodate in the reaction mixture in Experiment 2.

concentration = ... mol/dm^3 [1]

(ii) State which experiment, 1, 2, 3, 4 or 5, had the fastest rate of reaction. View solution

..... [1]

(g) Suggest why the volume of distilled water added to each experiment was increased as the volume of aqueous potassium iodate was decreased.

.....

..... [1]

(h) Give **one** change you could make to the apparatus used which would improve the results. Explain your answer.

View solution

change to apparatus .....

.....

explanation .....

.....

[2]

(i) Describe how the reliability of the results of this investigation could be checked.

View solution

.....

..... [1]

[Total: 19]

2 You are provided with two solids, solid **Q** and solid **R**.
Do the following tests on solid **Q** and solid **R**, recording all of your observations at each stage.

tests on solid Q

(a) Using a spatula, place approximately three quarters of solid **Q** in a boiling tube.
Add about 10 cm^3 of dilute sulfuric acid to the boiling tube. Test any gas produced.
Record your observations.

View solution

Keep the contents of the boiling tube for (c).

.....

.....

..... [3]

(b) Identify the gas given off in **(a)**.

View solution

..... [1]

(c) Filter the reaction mixture from **(a)**.
The filtrate is solution **S**. Pour 1 cm depth of solution **S** into a boiling tube.

View solution

(i) To solution **S** in the boiling tube add aqueous sodium hydroxide dropwise and then in excess.
Record your observations.

.....

.....

..... [2]

(ii) Explain why it is not possible to identify the cation contained in solution **S** from your observations in **(c)(i)**.

.....

..... [1]

(iii) Suggest an additional test that can be done on solution **S** that would allow you to identify the cation contained in solution **S**.

Do not carry out this test.

.....

..... [1]

(d) Place the remaining solid **Q** in a hard-glass test-tube. Heat the solid strongly until no further change is seen. Stop heating and leave the solid to cool.
Record your observations.

View solution

.....

.....

..... [2]

tests on solid R

(e) Carry out a flame test on solid **R**.
Record your observations.

View solution

..... [1]

(f) Place solid **R** in a boiling tube. Add about 10 cm^3 of distilled water to the boiling tube. Place a stopper in the boiling tube and shake the tube to dissolve solid **R** and form solution **R**.

View solution

Divide solution **R** into two approximately equal portions in two test-tubes.

(i) Add the first portion of solution **R** to the aqueous bromine provided in a test-tube.
Record your observations.

..... [1]

(ii) To the second portion of solution **R** add about 1 cm depth of dilute nitric acid followed by a few drops of aqueous silver nitrate.
Record your observations.

View solution

.....

..... [1]

(g) Identify solid **R**.

View solution

.....

..... [2]

[Total: 15]

3 Brass is a mixture of two metals, copper and zinc.

Copper does not react with dilute sulfuric acid. Zinc reacts with hot dilute sulfuric acid to form the soluble salt zinc sulfate.


View solution

Plan an investigation to find the percentage by mass of zinc in a sample of brass.
In your answer you should include how to calculate the percentage by mass of zinc.

You have access to normal laboratory apparatus.

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..... [6]

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Notes for use in qualitative analysis
Tests for anions

anion	test	test result
carbonate (CO_3^{2-})	add dilute acid	effervescence, carbon dioxide produced
chloride (Cl^-) [in solution]	acidify with dilute nitric acid, then add aqueous silver nitrate	white ppt.
bromide (Br^-) [in solution]	acidify with dilute nitric acid, then add aqueous silver nitrate	cream ppt.
iodide (I^-) [in solution]	acidify with dilute nitric acid, then add aqueous silver nitrate	yellow ppt.
nitrate (NO_3^-) [in solution]	add aqueous sodium hydroxide, then aluminium foil; warm carefully	ammonia produced
sulfate (SO_4^{2-}) [in solution]	acidify, then add aqueous barium nitrate	white ppt.
sulfite (SO_3^{2-})	add dilute hydrochloric acid, warm gently and test for the presence of sulfur dioxide	sulfur dioxide produced will turn acidified aqueous potassium manganate(VII) from purple to colourless

Tests for aqueous cations

cation	effect of aqueous sodium hydroxide	effect of aqueous ammonia
aluminium (Al^{3+})	white ppt., soluble in excess, giving a colourless solution	white ppt., insoluble in excess
ammonium (NH_4^+)	ammonia produced on warming	–
calcium (Ca^{2+})	white ppt., insoluble in excess	no ppt., or very slight white ppt.
chromium(III) (Cr^{3+})	green ppt., soluble in excess	grey-green ppt., insoluble in excess
copper(II) (Cu^{2+})	light blue ppt., insoluble in excess	light blue ppt., soluble in excess, giving a dark blue solution
iron(II) (Fe^{2+})	green ppt., insoluble in excess	green ppt., insoluble in excess
iron(III) (Fe^{3+})	red-brown ppt., insoluble in excess	red-brown ppt., insoluble in excess
zinc (Zn^{2+})	white ppt., soluble in excess, giving a colourless solution	white ppt., soluble in excess, giving a colourless solution

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Tests for gases

gas	test and test result
ammonia (NH_3)	turns damp red litmus paper blue
carbon dioxide (CO_2)	turns limewater milky
chlorine (Cl_2)	bleaches damp litmus paper
hydrogen (H_2)	'pops' with a lighted splint
oxygen (O_2)	relights a glowing splint
sulfur dioxide (SO_2)	turns acidified aqueous potassium manganate(VII) from purple to colourless

Flame tests for metal ions

metal ion	flame colour
lithium (Li^+)	red
sodium (Na^+)	yellow
potassium (K^+)	lilac
copper(II) (Cu^{2+})	blue-green

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